

Modern Chemistry Chapter Test B Answers

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[Modern Chemistry 24 Chapter Test . Name Class Date](#) | Chapter Test B, continued 7. The discovery of the nucleus was a result of Rutherford's observation that a small percentage of the positively charged particles bombarding the metal's surface a. were slightly deflected as they passed through the metal.

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[Modern Chemistry 141 Chapter Test Name Class Date Chapter Test B, continued](#) 18. When a strong acid is titrated with a weak base, the pH of the solution at the equivalence point is than 7. 19. A is a highly purified solid used to check the concentration of a standard solution. 20. A 1 M solution of NaOH will have a pH that is

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[Modern Chemistry 33 Chapter Test Name Class Date Chapter Test B, continued](#) 15. The energy state of an atom is called its ground state. 16. The number of waves that pass a point in one second is called. 17. When an electron drops from a higher-energy state to a lower-energy state, a(n) spectrum is produced. 18.

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[Modern Chemistry 91 Chapter Test Name Class Date Chapter Test B, continued](#) 27. Distinguish between ionic crystals and metallic crystals. PART V Write the answers to the following questions on the line to the left, and show your work in the space provided. [Modern Chemistry Chapter 12 Review Answer Key](#)

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[Modern Chemistry 20 Chapter Test Name Class Date Chapter Test A, continued](#) ____ 13. To calculate the number of atoms present in 2.0 mol of an element, you would a. add Avogadro's number of atoms per mole to 2.0 mole. b. subtract Avogadro's number of atoms per mole from 2.0 mole. c.

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[Holt McDougal Modern Chemistry 3 Chapter Test Chapter Test B, continued](#) 16 Modern chemistry chapter 3 test b answers. The measure of the ability of an atom in a chemical compound to attract electrons from another atom in the compound is called _____. 17. The energy required to remove one electron from an atom is called its _____. 18.

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CHAPTER 3 TEST continued Date Class FILL IN THE BLANK Write the correct term (or terms) in the space provided. 9. If a particular ompound is composed of elements A and B, the ratio of the mass of B to the mass of A will allw ys be the same. This is a statement of the law of exactly 12 g of carbon-12 is referred to as a(n) 11.

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(b) positive ions. (d) negative ions. 3. b Compared with the neutral atoms involved in the formation of an ionic compound, the crystal lattice that results is (a) higher in potential energy. (c) equal in potential energy. (b) lower in potential energy. (d) unstable. 4. b The lattice energy of compound A is greater in magnitude than that of ...

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b. Explain how an atom can exist in this state. Atoms consist of a positively charged nucleus, made up of protons and neutrons, that is surrounded by a negatively charged electron cloud. The positive and negative charges combine to form a net neutral charge. MODERN CHEMISTRY ATOMS: THE BUILDING BLOCKS OF MATTER 19

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[Modern Chemistry 137 Chapter Test Name Class Date Chapter Test A, continued](#) ____ 19. In the figure on the previous page, the pH at the equivalence point a. is equal to 7.0. b. is greater than 7.0. c. is less than 7.0. d. cannot be determined from the data given. ____ 20. In the figure on the previous page, the volume of titration standard

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CHAPTER 22 TEST Nuclear Chemistry Class MULTIPLE FHOICE On the line at the left of each statement, write the letter of the choice tha best completes the statement or answers the question. After converting units, the nuclear mass defect is equivalent to the a. atomic mass b. electrostatic force c. energy of chemical reaction

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CHAPTER 16 REVIEW Reaction Energy SECTION 2 SHORT ANSWER Answer the following questions in the space provided. 1. For the following examples, state whether the change in entropy favors the forward or reverse reaction: forward reaction a. HCl(l) ?? HCl(g) reverse reaction b. C 6H 12O 6(aq) ?? C 6H 12O 6(s) forward reaction c. 2NH 3(g) ...