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Problems Chemical

Chapter 10 Supplemental Problems Balance the following chemical equations. 1) $\text{SnS}_2(\text{s}) + \text{O}_2(\text{g}) \rightarrow \text{SnO}_2(\text{s}) + \text{SO}_2(\text{g})$ 2) $\text{C}_2\text{H}_6(\text{g}) + \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g})$ 3) $\text{Al}(\text{s}) + \text{HCl}(\text{aq}) \rightarrow \text{AlCl}_3(\text{aq}) + \text{H}_2(\text{g})$ 4) $\text{CoCO}_3(\text{s}) \rightarrow \text{CoO}(\text{s}) + \text{CO}_2(\text{g})$ Write a balanced equation for each of the following reactions, substituting symbols and formulas for names.

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2 1 Data Analysis Data Analysis 1. A sample of aluminum is placed in a 25-mL graduated cylinder containing 10.0 mL of water.

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Chapter 10 Supplemental Problems The Supplemental Problems for Chapter 10. Supplemental Problems for Chapter 10. 1) One overarching trend in atomic structures is that as the coordination number increases, the interatomic separation increases. Using the appropriate equations, answer the following questions.

Supplemental Problems for Chapter 10

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aluminum is placed in a 25-mL graduated cylinder containing 10.0 mL of water. The level of water rises to 18.0 mL.

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